

## Chapter11 Review Gases Answer Key

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### Chapter11 Review Gases Answer Key

CHAPTER 11 REVIEW Gases Class SHORT ANSWER Answer the following questions in the space provided. c c The molar mass of a gas at STP is the density of that gas (a) multiplied by the mass of 1 mol. (b) divided by the mass of 1 mol. nRT (c) multiplied by 22.4 L. (d) divided by 22.4 L. For the expression V = (a) increasing P (b) decreasing T

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Chapter 11 Review Gases Section 1 Answers CHAPTER 11 REVIEW . Gases . SHORT ANSWER . Answer the followin9 questions in the space provided. 1. c The molar mass of a gas at STP is the density of that gas (a) multiplied by the mass of 1 mol. (c) multiplied by 22.4 . L. (b) divided by the mass of 1 mol. (d) divided by 22.4 . L. 2. c For the expression V = n–T,

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CHAPTER . 11 . REVIEW . Gases . SHORT ANSWER . Answer the following questions in the space provided. 1. Consider the following data table: Approximate pressure (kPa) Altitude above sea level (km) 100 . o (sea level) 50 5.5 (peak of Mt. Kilimanjaro) 25 11 Get cruising altitude) < 0.1 22 (ozone layer) –.

### REVIEW Gases

Chapter 11 - Gases. Review Skills 11.1 Gases and Their Properties Ideal Gases Properties of Gases Discovering the Relationships Between Properties The Relationship Between Volume and Pressure. Internet: Boyle's Law.

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Modern Chemistry 97 Gases CHAPTER 11 REVIEW Gases SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. \_\_\_\_ The molar mass of a gas at STP is the density of that gas (a) multiplied by the mass of 1 mol. (c) multiplied by 22.4 L. (b) divided by the mass of 1 mol. (d) divided by 22.4 L. 2. \_\_\_\_ For the expression , P

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The volume of a fixed mass of gas is DIRECTLY proportional to the Kelvin Temperature, when pressure is held constant Equation: V1/T1 = V2/T2 P= Constant. Combined Gas Law. expresses the relationship between pressure, volume and temperature of a fixed amount of gas Equation: P1V1/T1 =P2V2/T2. The Ideal Gas Law.

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