

Chapter 8 Chemistry Covalent Bonding

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Chapter 8 Chemistry Covalent Bonding

1. In covalent bonds, electron sharing usually occurs so that atoms attain the electron configuration of noble gases. (B) 2. Atoms form double or triple covalent bonds if they can attain a noble gas structure by sharing two pairs or three pairs of electrons.

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a covalent bond between atoms in which the electrons are shared unequally a slightly negative charge After attracting electrons more strongly, the more-electronegative atom gains . . .

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Chemistry chapter 8 Covalent Bonding. STUDY. PLAY. Covalent bond. The chemical bond that results from sharing valence electrons. Molecule. Formed when two or more atoms bond covalently. Lewis structure. Represent the arrangement of electrons in a molecule. Sigma bonds. Single covalent bonds. Pi bond.

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Chemistry Chapter 8: Covalent Bonding. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. Brennallheart. Terms in this set (59) Covalent bond, a bond formed by the sharing of electrons between atoms. Molecule, a neutral group of atoms joined together by covalent bonds. Diatomic molecule.

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Single Covalent Bonds When only one pair of electrons is shared, such as in a hydrogen molecule, it is a single covalent bond. The shared electron pair is often referred to as the bonding pair. For a hydrogen molecule, shown in Figure 8.4,each covalently bonded atom equally attracts the pair of shared electrons.

Chapter 8: Covalent Bonding

A covalent bond in which the bonding electrons are most likely to be found in sausage-shaped regions above and below the bond axis of the bonded atoms Polar molecule A molecule in which one side of the molecule is slightly negative and the opposite side is slightly positive.

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Technically, any covalent bond between two different elements is polar. However, the degree of polarity is important. A covalent bond between two different elements may be so slightly imbalanced that the bond is, essentially, nonpolar. A bond may be so polar that an electron actually transfers from one atom to another, forming a true ionic bond.

Chapter 8 - Chemical Bonds - CHE 105/110 - Introduction to ...

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Pearson Chemistry Chapter 8 Covalent Bonding (Honors) covalent bond, molecule, diatomic molecule. HOFBrHCl, a bond formed by the sharing of electrons between atoms, a neutral group of atoms joined together by covalent bonds, a molecule consisting of two atoms.

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CHAPTER 8 SOLUTIONS MANUAL Covalent BondingCovalent Bonding Solutions Manual Chemistry: Matter and Change • Chapter 8 121 Section 8.1 The Covalent Bond pages 240-247 Practice Problems page 244 Draw the Lewis structure for each molecule. 1. PH 3 H H H— H H P respectively, for single, double, and triple P — 2. H 2 S H H H — H S S ...

Covalent BondingCovalent Bonding - Weebly

A covalent bond in which the bonding electrons are most likely to be found in sausage-shaped regions above and below the bond axis of the bonded atoms Polar molecule A molecule in which one side of the molecule is slightly negative and the opposite side is slightly positive.

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Covalent Bonding - Chapter 8 1. Covalent Bonding Or How I Learned to Love Sharing (But Remember, File Sharing is Illegal) 2. As you should remember, ionic compounds are solids at room temperatures that have one ion strip the electron(s) from the other elements' electron cloud. Some compounds do not give up their electrons so easily.

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8.3 Covalent Bonding •The majority of chemical substances do not have characteristics of ionic compounds. •A chemical bond formed by sharing a pair of electrons is called a covalent bond. •Both atoms acquire noble-gas electronic configurations.

AP Chemistry Chapter 8 Lecture Notes- Basic Bonding 8.1 ...

Section 8.2 The Nature of Covalent Bonding • OBJECTIVES: -Distinguish between a covalent bond and a coordinate covalent bond, and describe how the strength of a covalent bond is related to its bond dissociation energy.

Chapter 8 "Covalent Bonding"

Valence bond theory describes a covalent bond as the overlap of half-filled atomic orbitals (each containing a single electron) that yield a pair of electrons shared between the two bonded atoms. We say that orbitals on two different atoms overlap when a portion of one orbital and a portion of a second orbital occupy the same region of space.

8.1 Valence Bond Theory | Chemistry

Covalent bonding occurs when electrons are shared between atoms. Double and triple covalent bonds occur when four or six electrons are shared between two atoms, and they are indicated in Lewis structures by drawing two or three lines connecting one atom to another.

The Covalent Bond | Boundless Chemistry

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