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Solution . To get the requested answer, let us formulate a "stoichiometric" equation (molar quantities) for the reaction:  $C_2H_6 + 7O_2 \rightarrow 2CO_2 + 3H_2O$  Each  $C_2H_6$  (ethane) molecule requires 7 oxygen atoms for complete combustion. In molar quantities: 1 mole of  $C_2H_6 = 2 \times 12.01 + 6 \times 1.008 = 30.07$  g requires  $7 \times 15.9984$  g =  $1.12 \times 10^2$  g

### Session #2: Homework Solutions

Lesson #8 - Introduction to Logarithms Lesson #9 - Graphs of Logarithms Lesson #10 - Logarithm Laws Lesson #11 - Solving Exponential Equations Using Logarithms Lesson #12 - The Number e and the Natural Logarithm Lesson #13 - Compound Interest Lesson #14 - Newton's Law of Cooling

### COMMON CORE ALGEBRA II

1 Introduction 1 Motivation to Study Combustion 1 A Definition of Combustion 6 Combustion Modes and Flame Types 6 Approach to Our Study 8 References 8 2 Combustion and Thermochemistry 9 Overview 9 Review of Property Relations 9 Extensive and Intensive Properties 9 Equation of State , 10 Calorific Equations of State 11 Ideal-Gas Mixtures 13

### AN INTRODUCTION TO COMBUSTION Concepts and Applications

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### **An introduction to combustion : concepts and applications ...**

Updated January 09, 2020 A combustion reaction is a major class of chemical reactions, commonly referred to as "burning." In the most general sense, combustion involves a reaction between any combustible material and an oxidizer to form an oxidized product. It usually occurs when a hydrocarbon reacts with oxygen to produce carbon dioxide and water.

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